

Use with textbook pages 176-180.

Lewis diagrams

1. Define the following terms:

(a) Lewis diagram

(b) lone pair _____

(c) bonding pair _____

2. Draw Lewis diagrams for each of the following elements.

(a) boron

(b) nitrogen

(c) aluminium

(d) chlorine

3. Draw Lewis diagrams for each of the following ionic compounds.

(a) sodium oxide

(b) potassium chloride

(c) magnesium bromide

4. Draw Lewis diagrams for each of the following covalent compounds.

(a) carbon dioxide, CO_2 (b) phosphorus trifluoride, PF_3 (c) silicon tetrachloride, SiCl_4

5. Draw Lewis diagrams for each of the following diatomic molecules.

(a) chlorine, Cl_2

(b) nitrogen, N_2

(c) hydrogen, H_2

Name: _____

Date: _____

Block: _____

Chemical Reactions: Lesson 5 -Lewis Diagrams

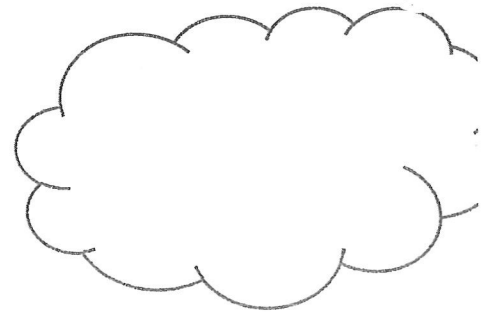
Lewis Diagrams: _____

Bohr Diagram	Magnesium	Lewis Dot Structure
	Atom	
	Ion	

Lewis Dot Structure

Ex. Oxygen Atom

Ex. Oxygen Ion



Ex. Negative Hydrogen Ion

Ex. Positive Hydrogen Ion

Ex. Aluminum Atom

Ex. Aluminum Ion

A: Lewis Diagram for IONIC COMPOUNDS

a) Draw a Lewis Diagram for Lithium Chloride.

Step 1:

Step 2:

b) Draw a Lewis Diagram for Calcium Iodide.

Step 1:

Step 2:

